

Curriculum Grade Book

Morgan County School District

Final, 01/11/2010

Chemistry Science

	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
Physical Science (25%)																														
<p>■ 1.1.1 (DOK 2) ASSESSED</p> <p>The learner will be able to classify or make generalizations about elements from data of observed patterns in atomic structure and/or position on the periodic table. (The periodic table is a consequence of the repeating pattern of outermost electrons.)</p>																														
<p>■ 1.1.2 (DOK) Supporting</p> <p>The learner will be able to understand that the atom's nucleus is composed of protons and neutrons that are much more massive than electrons; When an element has atoms that differ in the number of neutrons, these atoms are called different isotopes of the element.</p>																														
<p>■ 1.1.3 (DOK) Supporting</p> <p>The learner will be able to understand that solids, liquids, and gases differ in the distances between molecules or atoms and therefore the energy that binds them together. In solids, the structure is nearly rigid; in liquids, molecules or atoms move around each other but do not move apart; and in gases, molecules or atoms move almost independently of each other and are relatively far apart.</p>																														
<p>■ 1.1.4 (DOK) Supporting</p> <p>The learner will be able to understand that in conducting materials, electrons flow easily; whereas, in insulating materials, they can hardly flow at all. Semiconducting materials have intermediate behavior. At low temperatures, some materials become superconductors and offer no resistance to the flow of electrons.</p>																														
<p>■ 1.1.5 (DOK 2) ASSESSED</p>																														

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<p>The learner will be able to explain the role of intermolecular or intramolecular interactions on the physical properties (solubility, density, polarity, boiling/melting points) of compounds. The physical properties of compounds reflect the nature of the interactions among molecules. These interactions are determined by the structure of the molecular including the constituent atoms.</p>																														
<p>■ 1.1.6 (DOK 3) ASSESSED The learner will be able to identify variables that affect reaction rates; Predict effects of changes in variables (concentration, temperature, properties of reactants, surface area, and catalysts) based on evidence/data from chemical reactions. Rates of chemical reactions vary. Reaction rates depend on concentration, temperature, and properties of reactants. Catalysts speed up chemical reactions.</p>																														
<p>■ 1.1.7 (DOK 2) ASSESSED The learner will be able to construct diagrams to illustrate ionic or covalent bonding; Predict compound formation and bond type as either ionic or covalent (polar, nonpolar). Bonds between atoms are created when outer electrons are paired by being transferred (ionic) or shared (covalent). A compound is formed when two or more kinds of atoms bind together chemically.</p>																														
<p>■ 1.1.8 (DOK 3) ASSESSED The learner will be able to explain the importance of chemical reactions in a real world context; Justify conclusions using evidence/data from chemical reactions. Chemical reactions (e.g., acids and bases, oxidation,</p>																														

